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1) each element is made of atoms. 2) all atoms of a given element are identical there are  
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compound. Rutherford.

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need to understand before taking the exam.

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Basic Building Blocks of Matter. Atom, element, compound. Atom. An atom is the smallest unit of an element that maintains the chemical identity of that element. Element. An element is a pure substance that cannot be broken down into simpler, stable substances and is made of one type of atom. Compound.

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Molecules: the smallest unit of a substance that keeps all of the physical and chemical properties of that substance; it can consist of one atom or two or more atoms bounded together. (usually consists of two or more atoms combined in a definite ratio).

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The international standard unit of measure for the number of molecules in a substance is a mole. A mole is equal to Avogadro's number, or  $6.022 \times 10^{23}$ , which is standardized to the number of atoms that are present in 12 grams of Carbon-12.

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Chemistry A. What is Chemistry? -The study of the composition, structure and properties of matter and the reactions by which matter may be formed or converted into other forms. B. Classification of Matter Matter- an object that has mass and takes up space (volume) 1. Mixture- Two or more kinds of matter that can be separated by physical means.

Unit 1 Notes. Introduction to Chemistry - LPS

Unit 1 Study Guide. Learning Targets & Sample Questions (sample problems in italics) Measurement, Units, Problem Solving and Calculations: Review Chapter 1, Chemistry: An introduction and review Chapter 2, Measurements and Calculations. Be sure to review your notes, assignments and labs as you prepare for the upcoming test.

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